

pH-Logarithmen

pH = 0	$c(\text{H}_3\text{O}^+) = 10^{-0} \text{ mol/L}$	= 1 mol/L
pH = 1	$c(\text{H}_3\text{O}^+) = 10^{-1} \text{ mol/L}$	= 0.1 mol/L
pH = 2	$c(\text{H}_3\text{O}^+) = 10^{-2} \text{ mol/L}$	= 0.01 mol/L
pH = 3	$c(\text{H}_3\text{O}^+) = 10^{-3} \text{ mol/L}$	= 0.001 mol/L
pH = 4	$c(\text{H}_3\text{O}^+) = 10^{-4} \text{ mol/L}$	
pH = 5	$c(\text{H}_3\text{O}^+) = 10^{-5} \text{ mol/L}$	
pH = 6	$c(\text{H}_3\text{O}^+) = 10^{-6} \text{ mol/L}$	
pH = 7	$c(\text{H}_3\text{O}^+) = 10^{-7} \text{ mol/L}$	
pH = 8	$c(\text{H}_3\text{O}^+) = 10^{-8} \text{ mol/L}$	
pH = 9	$c(\text{H}_3\text{O}^+) = 10^{-9} \text{ mol/L}$	
pH = 10	$c(\text{H}_3\text{O}^+) = 10^{-10} \text{ mol/L}$	
pH = 11	$c(\text{H}_3\text{O}^+) = 10^{-11} \text{ mol/L}$	
pH = 12	$c(\text{H}_3\text{O}^+) = 10^{-12} \text{ mol/L}$	
pH = 13	$c(\text{H}_3\text{O}^+) = 10^{-13} \text{ mol/L}$	
pH = 14	$c(\text{H}_3\text{O}^+) = 10^{-14} \text{ mol/L}$	

Berechnen Sie die fehlenden pH-Werte mit dem TR:

pH = 0.6	$c(\text{H}_3\text{O}^+) = \underline{\hspace{2cm}}$	= $\underline{\hspace{2cm}}$ mol/L
pH = $\underline{\hspace{2cm}}$	$c(\text{H}_3\text{O}^+) = 10^{-3.1} \text{ mol/L}$	= $\underline{\hspace{2cm}}$ mol/L
pH = 3.5	$c(\text{H}_3\text{O}^+) = \underline{\hspace{2cm}}$	= $\underline{\hspace{2cm}}$ mol/L
pH = $\underline{\hspace{2cm}}$	$c(\text{H}_3\text{O}^+) = \underline{\hspace{2cm}}$ mol/L	= $1.58 \cdot 10^{-4} \text{ mol/L}$
pH = $\underline{\hspace{2cm}}$	$c(\text{OH}^-) = \underline{\hspace{2cm}}$ mol/L	= $1.58 \cdot 10^{-4} \text{ mol/L}$